

Chapter - Ionic bonding

Properties of Ionic compounds: —

(i) Ionic compounds have high melting point and boiling point as compared to covalent compounds.

(ii) They are generally soluble in water but insoluble in organic solvents.

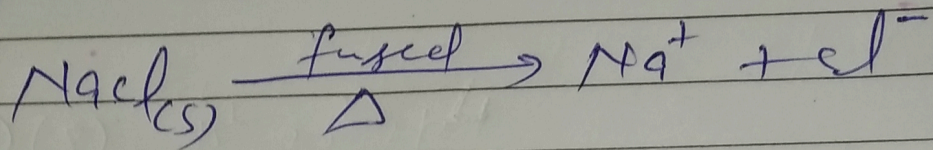
~~Ex~~ Exception — CaF_2 , CaCO_3 , MgCO_3 , BaCO_3 , SrCO_3 , CaSO_4 , SrSO_4 and BaSO_4 are insoluble in water.

(iii) They are generally solids at room temperature.

(iv) They ionise in aqueous solutions.

$$\text{NaCl}_{(s)} \xrightarrow{\text{water}} \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$$

(v) They ionise in fused state.



(vi) They are non-conductor of electricity in solid state. Because their ions are not mobile in solid state.

(vii) Their aqueous solutions are good conductors of electricity because their ions are mobile in aq.

(viii) They are good conductors of electricity in fused state because their ions are mobile in fused state also.